| Date Planned : __ / __ / __ | Daily Tutorial Sheet-9 | Expected Duration : 30 Min |
|---|---|---|
| Actual Date of Attempt : __ / __ / __ | Level-2 | Exact Duration :_____ |

**106.** Salt used for performing bead test in qualitative inorganic analysis is :

**(A)** $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O$ **(B)** $FeSO_4 \cdot (NH_4)_2SO_4 \cdot 6H_2O$

**(C)** $Na(NH_4)HPO_4 \cdot 4H_2O$ **(D)** $CuSO_4 \cdot 5H_2O$

**107.** Give the correct order of True (T) or False (F) for the following statements :

**I.** The solubility of alkali metal in liquid ammonia increases down the group.

**II.** Chlorine produces $OCl^-$ with hot NaOH solution and $ClO_3^-$ with cold NaOH solution.

**III.** NaCl is hygroscopic.

**IV.** The water of crystallisation in hydrated $BeCl_2$ is easily removed by heating.

**(A)** TFFF **(B)** TTTT

**(C)** TFFT **(D)** FTTT

**108.** Of the following statements only one is incorrect. The statement is :

**(A)** Calcium chloride decreases the freezing point of water

**(B)** The net material consumed in Solvay's process is a mixture of NaCl and $CaCO_3$

**(C)** $Na_2CO_3$ and $Ca(OH)_2$ are both used for water softening

**(D)** Alums form hexagonal crystals

**109.** The volume strength of 10 N $H_2O_2$ is :

**(A)** 112 **(B)** 11.2 **(C)** 0.112 **(D)** 56

**110.** Give the correct order of initials T or F for following statements. Use T if statement is true and F if it is false.

**I.** When lithium is burnt is oxygen, it forms superoxide $LiO_2$

**II.** Crude common salt is hygroscopic because of impurities of $CaSO_4$ and $MgSO_4$

**III.** Solubility of $CaI_2$ is more than that of $CaCl_2$

**IV.** A suspension of hydroxide of magnesium is used as a stomach antacid.

**(A)** FFFT **(B)** FFTT **(C)** TFFF **(D)** TTFF

**111.** The wire of flash bulb is made of :

**(A)** Mg **(B)** Na **(C)** Ca **(D)** Li

**112.** Which of the following is pearl ash ?

**(A)** $KMnO_4$ **(B)** $K_2CO_3$

**(C)** $K_2O_3$ **(D)** KOH

**113.** Industrially $H_2O_2$ is obtained from :

**(A)** 2-ethyl anthraquinol by oxidation and then reduction in a cyclic process

**(B)** $H_2SO_5$

**(C)** $H_2S_2O_8$

**(D)** $BaO_2$

**114.** Both Be and Al become passive on reaction with conc. nitric acid due to :

**(A)** The non-reactive nature of the metal

**(B)** The non-reactive nature of the acid

**(C)** The formation of an inert layer of oxide on the surface of the metals

**(D)** Formation of active layer of oxide on the surface of metals

**115.** In polymeric $(BeCl_2)_n$, there are :

**(A)** Three centre two-electron bonds

**(B)** Three centre four-electron bonds

**(C)** Two centre three-electron bonds

**(D)** Two centre four-electron bonds